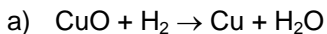
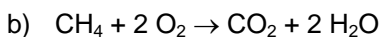


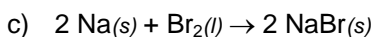
- 12** 6. For each of the following reactions, complete the summary table below the equation. If an element does not undergo any change, leave the last two columns blank. Also provide the formula of the oxidizing or reducing agent. An example is shown.



element	Initial Ox. No		Final Ox. No.	e ⁻ gained or lost	Oxidized or reduced	Oxidizing or Reducing Agent
Cu	+2	→	0	gain 2	reduced	oxidizing agent - CuO
O	-2	→	-2	0		
H	0	→	+1	lose 1	oxidized	reducing agent - H ₂



element	Initial Ox. No		Final Ox. No.	e ⁻ gained or lost	Oxidized or reduced	Agent
C	-4	→	+4	lose 8	oxidized	reducing agent - CH ₄
H	+1	→	+1	0		
O	0	→	-2	gain 2	reduction	oxidizing agent - O ₂



element	Initial Ox. No		Final Ox. No.	e ⁻ gained or lost	Oxidized or reduced	Agent
Na	0	→	+1	lose 1	oxidized	reducing agent Na
Br	0	→	-1	gain 1	reduced	oxidizing agent - Br ₂

- 5** 7. Not all of the following reactions are redox reactions. Place a check mark in the appropriate column for each reaction.

	Redox	Not Redox
a. $\text{Na}_2\text{S}(aq) + \text{FeCl}_2(aq) \rightarrow 2 \text{NaCl}(aq) + \text{FeS}(s)$	_____	_____ ✓
b. $2 \text{Na}(s) + 2 \text{H}_2\text{O}(l) \rightarrow 2 \text{NaOH}(aq) + \text{H}_2(g)$	_____ ✓	_____
c. $2 \text{KClO}_3(s) \rightarrow 2 \text{KCl}(s) + 3 \text{O}_2(g)$	_____ ✓	_____
d. $\text{SO}_2 + \text{H}_2\text{O} \rightarrow \text{H}_2\text{SO}_3$	_____	_____ ✓
e. $2 \text{Al} + 6 \text{HCl} \rightarrow 2 \text{AlCl}_3 + 3 \text{H}_2$	_____ ✓	_____