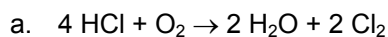


Unit 6: Redox Reactions and Electrochemistry

Practice Set 2: Recognizing Redox Reactions

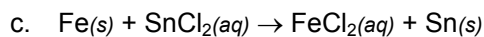
1. For each of the following reactions, complete the summary table below the equation. If an element does not undergo any change, leave the last two columns blank



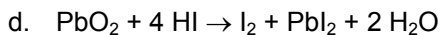
element	Initial Ox. No		Final Ox. No.	e ⁻ gained or lost	Oxidized or reduced	Agent
H		→				
Cl		→				
O		→				



element	Initial Ox. No		Final Ox. No.	e ⁻ gained or lost	Oxidized or reduced	Agent
Al		→				
O		→				

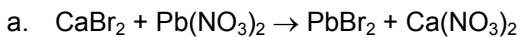


element	Initial Ox. No		Final Ox. No.	e ⁻ gained or lost	Oxidized or reduced	Agent
Fe		→				
Sn		→				
Cl		→				

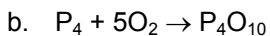


element	Initial Ox. No		Final Ox. No.	e ⁻ gained or lost	Oxidized or reduced	Agent
Pb		→				
O		→				
H		→				
I (to I ₂)		→				
I (to PbI ₂)		→				

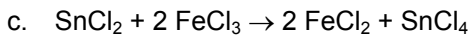
2. For each of these reactions, determine whether or not it is a redox reaction. If any are, identify oxidizing and reducing agents in those reactions.



element	Initial Ox. No		Final Ox. No.	e ⁻ gained or lost	Oxidized or reduced	Agent
		→				
		→				



element	Initial Ox. No		Final Ox. No.	e ⁻ gained or lost	Oxidized or reduced	Agent
		→				
		→				



element	Initial Ox. No		Final Ox. No.	e ⁻ gained or lost	Oxidized or reduced	Agent
		→				
		→				