Chemistry 30

Unit 6: Redox Reactions and Electrochemistry

Practice Set 5: Electrolytic Cells

1. Using either an activity series (Table 11.2) or Table of Standard Reduction Potentials, predict whether a reaction takes place, and if so, give a balanced reaction equation.

a.
$$Ag(s) + HCI(aq) \rightarrow$$

b.
$$Mg(s) + FeSO_4(aq) \rightarrow$$

c.
$$Cu(s) + AuCl_3(aq) \rightarrow$$

d.	$Sn(s) + Al_2(SO_4)$	$)_3$ (aq) \rightarrow
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- 2 Consider the electrolysis of water. Describe the events at the anode in terms of:
 - a. the reaction
 - b. pH
 - c. gas produced