

SOLUBILITY PRODUCT CONSTANTS

FOR VARIOUS COMPOUNDS AT 25°

* at 18°C ** at 27°C

Name	Formula	K_{sp}	Name	Formula	K_{sp}
Aluminum hydroxide	Al(OH) ₃	3.7×10^{-15}	Lithium carbonate	Li ₂ CO ₃	1.7×10^{-3}
Barium carbonate	BaCO ₃	2.6×10^{-9}	Magnesium carbonate	MgCO ₃	6.8×10^{-6}
Barium chromate	BaCrO ₄	1.2×10^{-10}	Magnesium fluoride	MgF ₂	$6.4 \times 10^{-9**}$
Barium fluoride	BaF ₂	1.7×10^{-6}	Magnesium hydroxide	Mg(OH) ₂	5.6×10^{-12}
Barium sulfate	BaSO ₄	1.1×10^{-10}	Mercury(II) sulfide	HgS	3.0×10^{-54}
Calcium carbonate	Ca CO ₃	8.7×10^{-9}	Silver bromate	AgBrO ₃	5.3×10^{-5}
Calcium fluoride	CaF ₂	3.9×10^{-11}	Silver bromide	AgBr	5.4×10^{-13}
Calcium oxalate	CaC ₂ O ₄	2.3×10^{-9}	Silver carbonate	Ag ₂ CO ₃	8.5×10^{-12}
Calcium sulfate	CaSO ₄	2.0×10^{-4}	Silver chloride	AgCl	1.8×10^{-10}
Copper(I) iodide	CuI	1.3×10^{-12}	Silver chromate	AgCrO ₄	1.1×10^{-12}
Copper(II) iodate	Cu(IO ₃) ₂	6.9×10^{-8}	Silver iodide	AgI	8.5×10^{-17}
Copper(II) oxalate	CuC ₂ O ₄	2.9×10^{-9}	Silver sulfate	Ag ₂ SO ₄	1.2×10^{-5}
Copper(II) sulfide	CuS	1.0×10^{-44}	Silver sulfide	SrS	1.8×10^{-50}
Iron(II) hydroxide	Fe(OH) ₂	4.9×10^{-17}	Strontium carbonate	SrCO ₃	5.6×10^{-10}
Iron(III) hydroxide	Fe(OH) ₃	2.6×10^{-39}	Strontium fluoride	SrF ₂	4.3×10^{-9}
Iron(II) sulfide	PbS	$3.7 \times 10^{-36*}$	Strontium sulfate	Sr SO ₄	3.4×10^{-7}
Lead(II) bromide	PbBr ₂	6.6×10^{-6}	Thallium(I) bromide	TlBr	3.8×10^{-6}
Lead(II) chloride	PbCl ₂	1.2×10^{-5}	Thallium(I) chloride	TlCl	2.2×10^{-4}
Lead(II) iodide	PbI ₂	8.5×10^{-9}	Thallium(I) iodide	TlI	3.1×10^{-8}
Lead(II) sulfate	FeSO ₄	1.8×10^{-8}	Zinc sulfide	ZnS	2.0×10^{-25}